Do Now

3/19

$NaHCO_3 + CH_3COOH \rightarrow CO_2 + H_2O + NaCH_3COO$

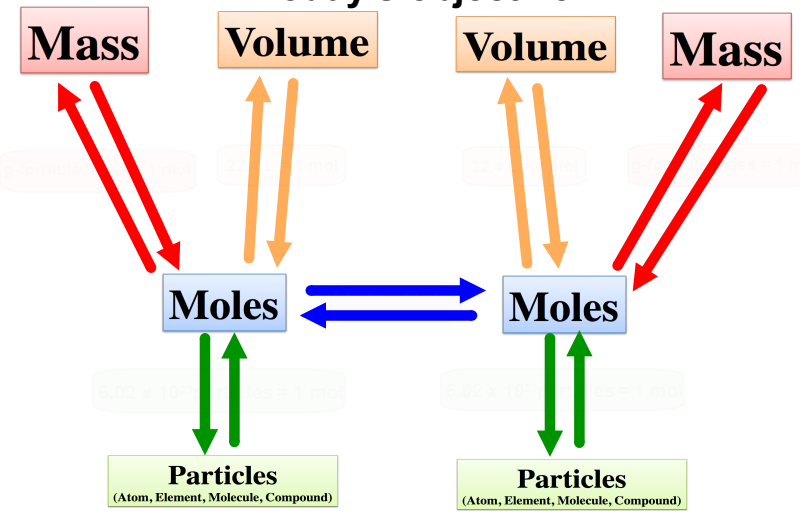
Moles					
Particles					
Mass					
Volume					

Fill out this chart using what you are given and knowledge about balancing equations.

5 mL of CH_3COOH (vinager)

g of $NaHCO_3$ (Baking soda)

Today's objective



Molar Mass

$NaHCO_3$

Element	Atomic Mass	#	TOTAL
C	12.011 g/mol	x 1	12.011 g/mol
H	1.008 g/mol	x 1	1.008 g/mol
O	15.999 g/mol	x 3	47.997 g/mol
Na	22.990 g/mol	x 1	22.990 g/mol

$$84.006 \ \frac{g}{mol}$$

$NaHCO_3$

grams $NaHCO_3$ = ? mL of CO_2

$$\frac{84.006 \ g}{1 \ mol} \ NaHCO_3$$

$$\frac{g_{NaHCO_3}}{1} \times \frac{1 \text{ mol}}{84.006 \text{ g}} =$$

Individual Exit Ticket

7.2.b) Explain why mole ratios are central to solving stoichiometry problems.

1 mol = 6.02 x 10^{23} particles
1 mol = Molar Mass (g-formula-mass)
1 mol = 22.4 L (gas at STP)